

Electrochemical Determination of Hydrogen Entry to HSLA Steel **during** Pickling

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Pickling with hydrochloric acid is a standard method to clean steel surfaces before hot-dip galvanizing. When normal low strength steels are pickled, hydrogen formed in pickling reactions does not have any significant harmful effect on the mechanical properties of steel. However, in pickling of steels with higher strength, the penetration of hydrogen into the steel may cause severe damages. **The effect of pickling of high strength low alloy (HSLA) steels was investigated using a cell construction based on Devanathan-Stachurski method with modified anodic surface treatment and hydrogen production using acid.** The penetration and the permeability of hydrogen were measured using an electrochemical cell with hydrochloric acid on the one side of the steel sample and a solution of NaOH on the other side. No protective coating, e.g. palladium on the anodic side of the sample is **needed**. The penetration rate of hydrogen into the steel and exit rate from the steel were lower for higher strength steel.

KEYWORDS: hydrogen permeation, high strength low alloy steel, pickling

1 INTRODUCTION

The number and the applications of new high strength low alloy (HSLA) steels with enhanced formability are constantly increasing. It is generally accepted that the transition from mild steel to HSLA steel occurs at yield strength of about 300 MPa. These steels have the potential to introduce weight savings while improving performance. Depending on design, the higher strength can evolve into better fatigue and crash performance, while maintaining or even reducing required material thickness [1]. However, HSLA steels may also give rise to some risks. For example, under certain circumstances hydrogen introduced into steel during its manufacture, subsequent component fabrication, or in service may result in brittle failures at applied stresses far below the yield strength or the nominal design stress for the materials [2].

In this paper, we have studied hydrogen intrusion to HSLA steels during acid pickling to remove surface films before hot-dip galvanizing. The electrolytically produced hydrogen that penetrates **into** metal lattice can originate from corrosion, pickling, or electroplating [3]. In the pretreatment of hot-dip galvanizing process, hydrogen may be absorbed in the steel during the pickling stage through contact with the hydrochloric acid which oxidizes the metal and forms adsorbed hydrogen that partly enters the metal lattice. After hot-dip galvanizing the zinc layer prevents or delays the exit of the hydrogen from the steel [4]. Steels with tensile strength at the level of 800-1000 MPa or higher are considered to be the most susceptible ones to hydrogen embrittlement [5,6]. The majority of steels have generally tensile strength in the range of 200-450 MPa, and they are apparently not subject to hydrogen embrittlement problems. Special consideration must be given to high strength steels if they are required to be pickled [3].

Inhibitors are used to prevent pitting of steel surface in case of over-pickling. They also minimize the dissolution of iron, decrease the acid consumption and lower the formation of harmful fumes. Normally, inhibitors provide good protection against hydrogen absorption during acid cleaning. It is known that inhibitors, such as benzotriazole obstruct or decrease the diffusion of

atomic hydrogen into the steel. In addition, some nitrogen bearing organic compounds such as imidazoline, have been found to be highly effective inhibitors of hydrogen penetration into low carbon steels. [2, 7].

The Devanathan-Stachurski method is the most frequently utilized technique to measure the hydrogen permeation in metals. Hydrogen is produced on the cathodic side of the double cell using cathodic polarisation. Part of this hydrogen diffuses through the sample and on the anodic side exiting hydrogen atoms are oxidized by applying an anodic current. Thickness of steel samples in hydrogen permeation experiment is typically 0.3–1.0 mm. Galvanostatic current to produce hydrogen has varied from 0.05-100 mA/cm² [8-10]. The aim of this investigation was to measure the permeation of hydrogen through HSLA steels using the Devanathan-Stachurski method with modified anodic surface treatment and hydrogen production using acid. The permeation is commonly studied using external polarization to produce the hydrogen, but we used the pickling acid as the hydrogen source. The samples had no Pd coating on the exit surface, but the exit surface was passivated potentiostatically using NaOH before permeation measurements. Passivation was done to produce a steady background current on the steel. A schematic diagram on the reactions and phenomena during the measurements is shown in Figure 1. In the anodic compartment steel is first dissolved and passivated at constant potential. The cathodic compartment is empty during the passivation stage. During the permeation tests in the cathodic compartment steel is corroded by acid at the free corrosion potential and part of the produced hydrogen enters the steel diffusing to the exit surface in the anode compartment. When hydrogen diffuses through the steel also hydrogen evolution begins on the anodic surface and this is detected as current increase.

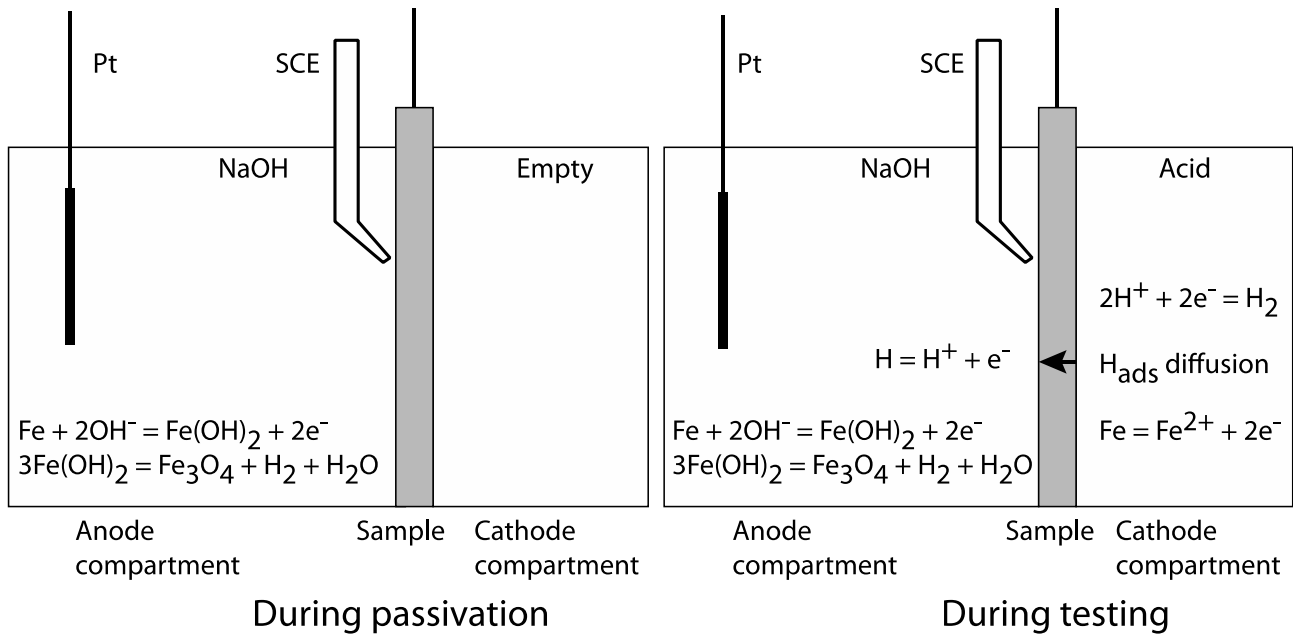


Figure 1. Schematic setup of the test cell with anodic and cathodic reactions at passivation and permeation test stages. Pt is platinum counter electrode and SCE is saturated calomel reference electrode.

Usually the anodic side of the steel sample is coated with palladium to ensure a uniform surface and to minimise the anodic dissolution of the metal during permeation current measurements. Palladium film favours the kinetics of hydrogen oxidation. Without Pd layer the oxidation of hydrogen atoms could be incomplete due to the presence of an oxide film that could build up on a metal sample when anodically polarised in an alkaline medium. The oxide film acts as a barrier against hydrogen release. According to Manolatos et al. [11] without a palladium coating on the exit side of the sample the surface phenomena are not controlling and the passive layer on the exit side changes with time preventing stabilization of the hydrogen concentration on the exit side and therefore stationary conditions cannot be obtained. This is the reason why most of the investigations have been carried out using palladium coating. The variation in hydrogen concentration is observed as non-steady state of the charging current curve. However, it is possible to obtain reproducible results by controlling the parameters related to the formation of the passive

layer [11]. Our approach was to use long passivation times and subtract the steady passive current from hydrogen oxidation transient currents.

The charging of steel with hydrogen is usually done electrolytically or in hydrogen-containing atmosphere, c.f. [12], and the gaseous charging and electrolytic charging are considered equivalent. Identical setup with freely corroding charging side has been used for example in [13] to study the effect of pickling inhibitors for hydrochloric acid, and in [14, 15] to study corrosion of pipeline steels. In [13] the effect of inhibitors on hydrogen permeation was calculated by comparing permeation current with and without inhibitor. In [14] the permeation current was used to evaluate the effect of additives and corrosion films on hydrogen absorption. In [15] the permeation current increased when the H_2S concentration in hydrogen entry environment increased. These examples indicate that use of freely corroding surface can be used to charge steel with hydrogen. In our work, the permeation of hydrogen was investigated using 1-4 mm thick steel samples, and several analysis methods were applied for estimate the diffusion coefficients.

2 EXPERIMENTAL

A double-compartment test cell was constructed to measure the permeation of hydrogen through steel (Fig. 2). Material of this double cell was Teflon and the steel sample was installed between the two separate cells. The electrochemical tests were done using Autolab PGSTAT30 potentiostat with Autolab Software 4.9.

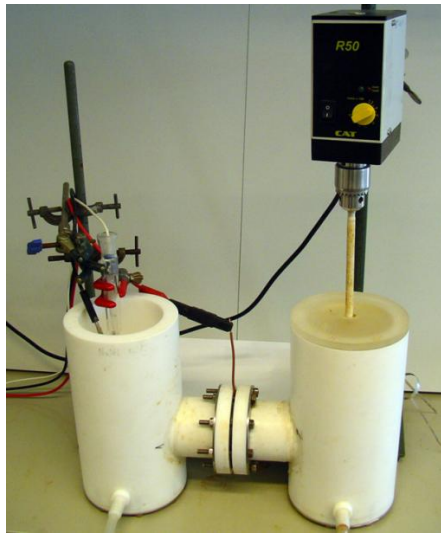


Fig. 2. Test cell for permeation experiments, with compartment with pickling solution on the right side and 0.2 M NaOH on the left side.

Test materials were high strength **low-alloy** (HSLA) steels with yield strength of 355, 500, 700 and 900 MPa. Samples for experiments were cut from rectangular hollow sections 50 mm wide with 6 mm material thickness. The material in these hollow sections is hardened, hot rolled steel and corresponding grade carbon steel S235 was used as reference material. The tensile strength of this steel is 235 MPa. Chemical compositions of test materials are shown in Table 1. CEV (Carbon Equivalent for estimation of steel weldability) values calculated using $CEV = C + Mn/6 + (Cr+Mo+V)/5 + (Ni+Cu)/15$ were taken for small material thicknesses used in this study. The analyses and material properties were supplied by the manufacturer and refer to general product qualities.

Table 1. Chemical compositions of steels used in experiments.

	C _{max}	Si	Mn _{max}	P _{max}	S _{max}	Al _{min}	CEV _{max}
S355	0.18	max 0.5	1.60	0.025	0.02		0.39

S500	0.18	max 0.5	1.70	0.020	0.015	0.02	0.43
S700	0.1	0.15- 0.25	2.0	0.020	0.01	0.015	0.41
S900	0.15	max 0.30	1.25	0.020	0.010		0.57
S235	0.17		1.40	0.035	0.035		0.35

The surface area of the sample that was in contact with solution was 19.6 cm². The thickness of the sample area exposed to the solutions was reduced by turning (Fig. 3). The thickness of exposed sample area varied from 1 to 4 mm. After thickness adjustment, the anodic surface of the sample was ground with emery paper (grade 400) followed by washing in citric acid just before electrochemical experiments. No coating was used on the anodic side of the sample.

The anodic potential in potentiostatic permeation experiments was determined by potentiodynamic polarisation measurements for carbon steel S235 and S700 in 0.2 M NaOH solution at room temperature. Anodic polarisation curve from the open circuit was measured with a potential sweep rate 100 mV/min.



Fig. 3. Test sample after turning. The middle part is for hydrogen permeation and the machined side was in the hydrogen evolving side.

Hydrochloric acid diluted to 16.5% (5.4 M) was used in the permeation measurements to produce hydrogen on steel surface. The acid was used both without dissolved iron and with Fe^{2+} concentrations 14 g/L (0.25 M) and 45 g/L (0.8 M) to simulate a used pickling bath. Iron was added as $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$. Temperature of the solution was 20 °C. The sample was installed between the two compartments of the cell and 1 L of 0.2 M NaOH solution was added into the anodic side of the test cell. Thereafter sample was polarized to a potential of 0 mV vs. SCE to form a passive layer on the surface and the current was measured up to steady state situation (about 15 h). The passive film in alkaline solution is usually magnetite [16,17]. The passivation procedure causes initially an increase in the current due to the corrosion of the steel when the oxide layer is formed and this current decreases with time as the passive film reaches a steady state. In the passive state steel dissolves, but at a low and constant rate. When the passive film had reached a steady state, hydrochloric acid (1 L) was introduced in the cathodic compartment. When hydrogen started to diffuse to the exit side a current increase was seen. When the current reached a steady maximum level, the acid was removed through the bottom valve and after some time a decrease of the anodic current was observed. During the whole measurement, the anodic side of the sample was kept at constant potential. In the beginning of the test the current was due to anodic oxidation and passivation of steel. Oxidation of hydrogen that had diffused to exit side was seen as current increase. When hydrogen was stopped by removing the acid current started to decrease and reached finally the same passive current level as before the introduction of acid. This background current is subtracted from the current measured during the permeation measurements. Thus, the effect of only hydrogen diffusion on the current could be observed.

To increase the uptake of hydrogen into the metal, recombination poisons such as arsenic, hydrogen sulphide and thiourea can be used [18]. These elements block adsorption sites on electrode surface diminishing the coverage by adsorbed hydrogen. The adsorbed H is an intermediate resulting from discharge of H^+ and its recombination produces H_2 in the cathodic

hydrogen evolution. No tests were done with hydrogen poisons as they are not used in pickling baths. The use of recombination poisons was not found to be necessary as detectable amount of hydrogen was generated from the concentrated hydrogen chloride solutions used in this work.

3 RESULTS AND DISCUSSION

The open circuit potential of carbon steel S253 was -490 mV vs. SCE and a wide passive range up to the potential of +400 mV vs. SCE was found in 0.2 M NaOH-solution (Fig. 4). Based on these results the potential of 0 mV vs. SCE was chosen for the permeation measurements. The selected potential is on the passive range of tested steels and it is higher than the equilibrium potential of H^+/H_2 reaction so oxidation of diffused hydrogen was guaranteed. The current density at this potential is less than $0.1 \mu A/cm^2$. The dissolution rate of steel is at very low level, approximately $1 \mu m/year$.

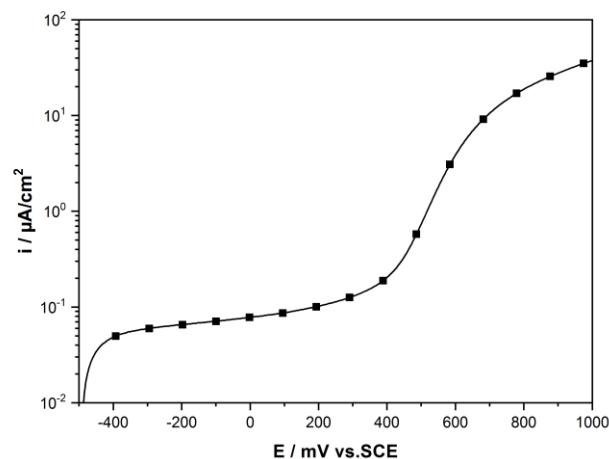


Fig. 4. The polarisation curve of carbon steel S253 in 0.2 M NaOH at 20 °C.

A typical current density vs. time curve in an experiment on steel S355 is shown in Fig. 5. The current is related only to the reaction rates on steel surface in the anodic compartment containing NaOH. A high current density is observed at the beginning of the experiment, when the steel

surface is still in active state, and much lower current density of about $0.3 \mu\text{A}/\text{cm}^2$ was reached when the steel passivates. Hydrochloric acid was added on the cathodic side of the cell after 18 h when the anodic surface had reached the passive state. After a delay of over two hours an increase in the current was seen in the anodic compartment. This current increase is due to hydrogen that had diffused through the steel and was oxidized on the anodic side. A steady maximum was observed after about 5 h from the acid addition indicating steady flux of hydrogen through the steel sample. At this point the hydrochloric acid was removed from the cathodic side of the cell. The anodic current density started to decrease as no more hydrogen was produced on the acid side and hydrogen charged into the steel was consumed. At the end of the test the current was approximately same passive current as before acid addition.

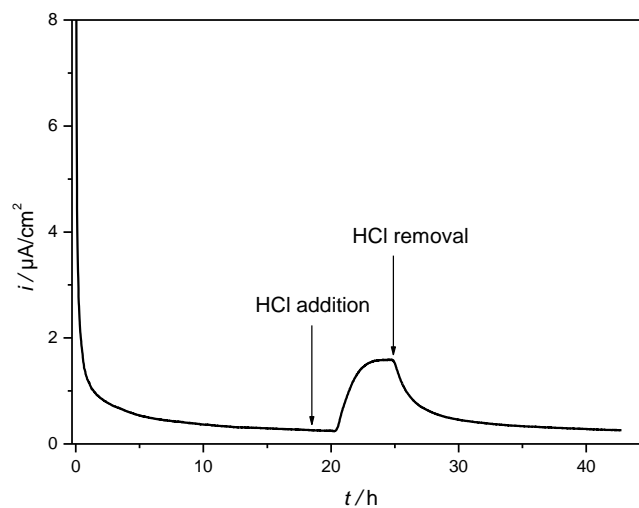


Fig. 5. A typical current density vs. time curve measurement in passivation and permeation experiment in 16.5% HCl for steel S355.

Results for hydrogen permeation experiments for high strength steel are shown in **Fig. 6**. The curves show only the part after adding HCl in the cathodic compartment. The hydrogen penetration time through the sample is similar for all steels with similar thickness, as shown in **Fig. 6**. This is seen as an increase in the current density after HCl addition. The maximum current density is lowest for the steel S900 and highest for steel S500. The grade of the steel does not have direct

correlation with the maximum current. The maximum current density gives the maximum penetration rate of hydrogen through the steel. Theoretically integration of current vs. time curve gives the amount of hydrogen penetrated through the sample.

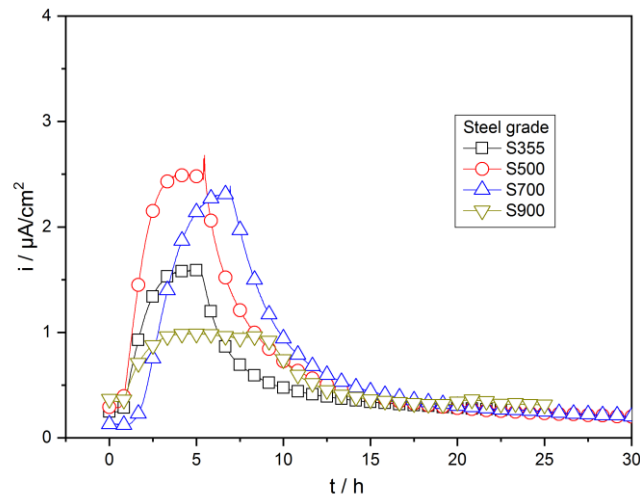


Fig. 6. Permeation of hydrogen in 16.5% HCl for different steel grades with sample thickness 1 mm.

Introduction of iron into the pickling acid has only a minor effect on the current vs. time curve (Fig. 7). As an example, the influence of ferrous ion concentrations 0 g/L, 14 g/L and 45 g/L on hydrogen entry for S700 steel is shown in Fig. 7. Dissolved iron had only a minor effect on the permeation of hydrogen in the studied steels.

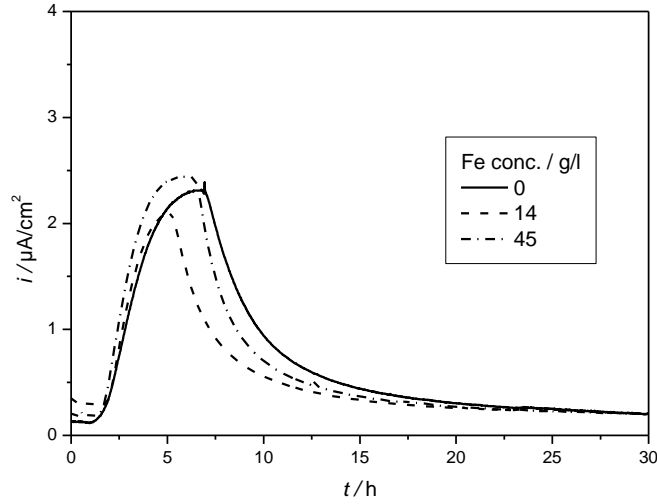


Fig. 7. The influence of iron content of the HCl on the permeation in steel S700.

The hydrogen diffusion coefficient was calculated by using four different methods given in [19, 20]:

- (1) The time lag method
- (2) Breakthrough time
- (3) Half-rise time
- (4) Analysis of the beginning decay transient

The analysis methods are based on the solution of the diffusion equation with constant concentration boundary conditions. The solution for the charging transient, derived by separation of variables is shown in equation (1)

$$I = nFAD \frac{c_1}{L} \left[1 + 2 \sum_{n=1}^{\infty} (-1)^n \exp \left(- \frac{D n^2 \pi^2 t}{L^2} \right) \right] \quad (1)$$

and for the decay transient (both ends at zero concentration and linear initial condition) the following solution is obtained, equation (2)

$$I = nFAD \frac{c_1}{L} 2 \sum_{n=1}^{\infty} (-1)^{n+1} \exp \left(- \frac{D n^2 \pi^2 t}{L^2} \right) \quad (2)$$

Fig. 8 shows the principles for determination of characteristic times for calculation of diffusion coefficients.

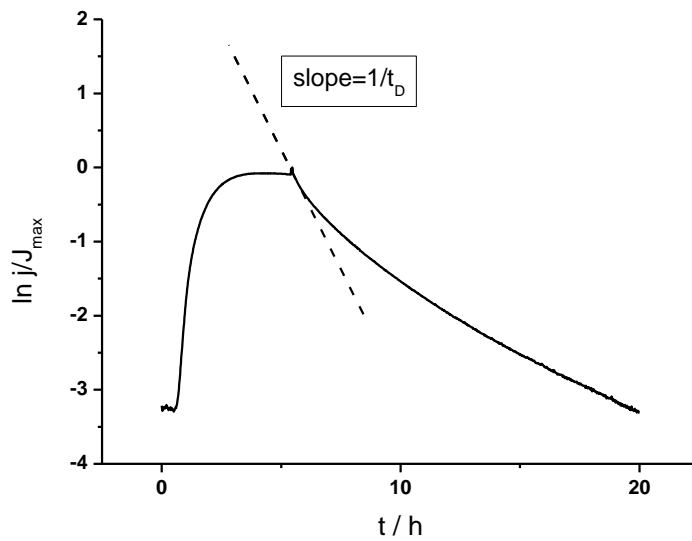
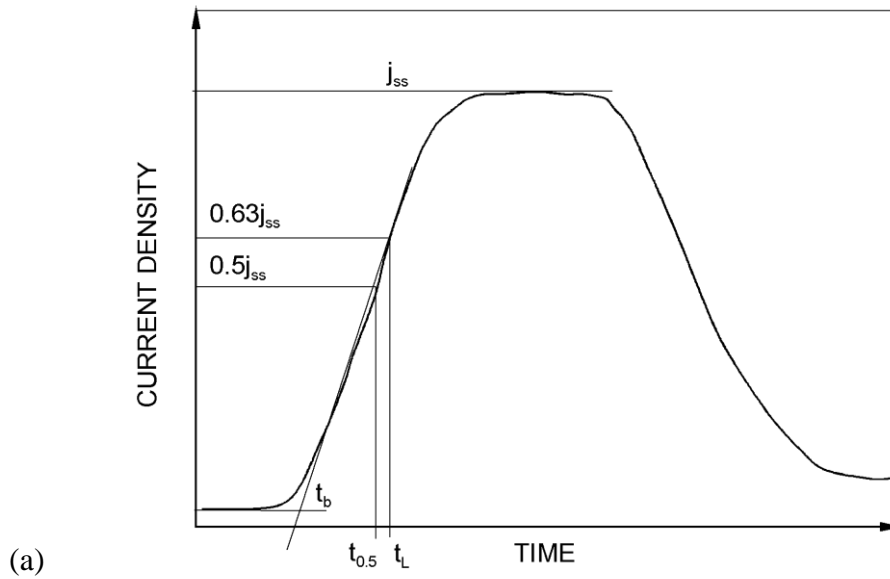


Fig. 8. Determination of characteristic times for calculation of diffusion coefficients. (a) time lag t_D , Breakthrough time t_b and Half-rise time $t_{0.5}$, (b) Decay time constant t_D .

In the Time lag method analysis is done by integrating the rising curve to calculate the quantity of hydrogen, which has permeated through the sample as a function of time. An extrapolation of the plot of quantity against time gives the time lag t_L which is related to the diffusion constant by (3):

$$D_L = \frac{1}{6} \cdot \frac{L^2}{\pi^2 \cdot t_L} \quad (3)$$

The t_L in equation (3) can be obtained by determining the time at which the rate of permeation is 0.63 times the steady-state value described by current density j_{ss} . [19].

In the Breakthrough time method, the time t_b is obtained by drawing tangent in the rising part of the curve. The tangent crosses the time axis at point t_b , equation (4) [19].

$$D_B = 0,5 \cdot \frac{L^2}{\pi^2 \cdot t_b} \quad (4)$$

In the Half-rise time method, the characteristic time $t_{0.5}$ corresponds to the current density that is 0.5 times the current density at steady-state situation j_{ss} . Diffusion coefficient can then be calculated by equation (5). [19]

$$D_{0.5} = 0,138 \cdot \frac{L^2}{t_{0.5}} \quad (5)$$

In the Decay transient analysis, a characteristic time constant t_D is determined using the slope of current decay after the steady state. The obtained current versus time is corrected for offsets and divided by its maximum value at the steady state. Taking the logarithm of the obtained normalized current one obtains for the beginning decay transient a linear section from which the slope is the inverse decay time constant, equation (6) [20].

$$D_D = \frac{1}{t_D} \frac{L^2}{\pi^2} \quad (6)$$

When the sample thickness increases the diffusion coefficient value in replicate tests seem to become more reproducible. At sample thicknesses of 1 mm the diffusion coefficient values were from 10^{-9} to 10^{-6} cm²/s. When the sample thickness was 1.5-2 mm the range was from 10^{-8} to 10^{-6} cm²/s and with sample thicknesses of 4 mm from 10^{-7} to 10^{-6} cm²/s. This effect was clearer with 500 MPa and 700 MPa samples. This effect agrees with the finding of Charca et. al [21] by using Armco iron. It is generally easier to attain solubility limits with thinner metallic samples than thicker ones, which contain more material defects and hence trapping sites.

The diffusion coefficient was not dependent on the analysis method. The coefficient decreased slightly with increasing strength. The results presented in Table 2 show that both charging and discharging transients have similar diffusion coefficients. This is important for the argumentation, whether deep traps are significantly involved and filled during the charging.

Table 2. Average values of diffusion coefficients (cm²/s) for test materials calculated using the four methods.

Method/steel	355 MPa	500 MPa	700 MPa	900 MPa
$D_{0.5}$	$1.82 \cdot 10^{-7}$	$1.6 \cdot 10^{-7}$	$1.61 \cdot 10^{-7}$	$2.5 \cdot 10^{-7}$
D_L	$1.16 \cdot 10^{-7}$	$8.78 \cdot 10^{-8}$	$3.2 \cdot 10^{-8}$	$2.54 \cdot 10^{-8}$
D_B	$1.96 \cdot 10^{-7}$	$1.49 \cdot 10^{-7}$	$1.73 \cdot 10^{-7}$	$1.93 \cdot 10^{-7}$
D_D	$1.76 \cdot 10^{-7}$	$1.81 \cdot 10^{-7}$	$9.76 \cdot 10^{-8}$	$1.26 \cdot 10^{-7}$

Filling of traps would lead to an internal accumulation of hydrogen, and this hydrogen that cannot be detected in the charging transient anymore. Since deep traps are preferentially filled

before the hydrogen can diffuse on, the penetration of the hydrogen is delayed, which would show up in the analysis as changed diffusion coefficients. For the decay transient, when all the deep traps were already filled, the diffusion of the hydrogen would proceed normally. Thus, whenever there would be deep traps present, this should be visible as a significant difference in the diffusion coefficient determination methods for charging (equation 3-5) compared to the discharge method (equation 6). That the diffusion coefficients obtained from the decay transient are similar to the other method, shows thus, that an involvement of traps for the tested steels can be neglected. That iron based samples do not have a significant **number** of traps has been already mentioned in [22].

To evaluate a concrete problem case, the here obtained diffusion coefficients can be used to simulate the penetration of hydrogen during pickling and the release of it during the time between pickling and hot dip galvanizing. This, however, is very much depending on the dimensions of the steel product and on the degree of deformation within that product. All this we leave for future investigations.

4 CONCLUSIONS

The purpose of this paper was to study the intrusion of hydrogen during pickling of steel. Hydrogen permeation **can be caused** by pickling acid, **such** as HCl, instead of cathodic hydrogen production that is mainly used in investigations found in the literature.

No protective coating, e.g. palladium on the anodic side of the sample is required. Permeation measurement needs to be carried out after passivation of steel and this low background current is to be subtracted from the measured current before the calculation of the diffusion coefficient.

The permeation rate of hydrogen into the steel and desorption rate from the steel are slightly lower for higher strength steel. This means that if these steels are plated soon after pickling, for example in hot-dip galvanizing, hydrogen has no time to leave steel and hydrogen induced failures may occur.

Iron dissolved in pickling acid has no essential effect on the permeation of hydrogen.

The authors declare that there is no conflict of interest regarding the publication of this paper.

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